

**Percent Yield WS adapted from *ChemFiesta.com***

- 1)
  - a) Write the balanced equation:  
lithium hydroxide + potassium chloride  $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
  - b) The reaction began with 20.0 grams of lithium hydroxide. What is the theoretical yield of lithium chloride?
  - c) The reaction produced 6.00 grams of lithium chloride. What is the percent yield?
  
- 2)
  - a) Write the balanced equation:  
beryllium + hydrochloric acid  $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
  - b) A student's theoretical yield of beryllium chloride was 10.7 grams. If the actual yield was 4.50 grams, what was the percent yield?
  
- 3)
  - a) Write the balanced equation:  
sodium chloride + calcium oxide  $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
  - b) What is the theoretical yield of sodium oxide if a chemist begins with 20.0 grams of calcium oxide?
  
- 4)
  - a) Write the balanced equation:  
iron(II) bromide + potassium chloride  $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
  - b) What is the theoretical yield of iron (II) chloride if the reaction begins with 34 grams of iron (II) bromide?
  - c) What is the percent yield of iron (II) chloride if my actual yield is 4.0 grams?
  
- 5)
  - a) Balance the following equation:  $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
  - b) If a reaction starts with 5.0 grams of  $\text{C}_3\text{H}_8$ , what is the theoretical yield of water?
  - c) If the percent yield of 75%, how many grams of water were produced?