

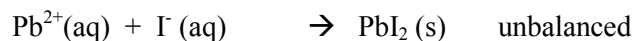
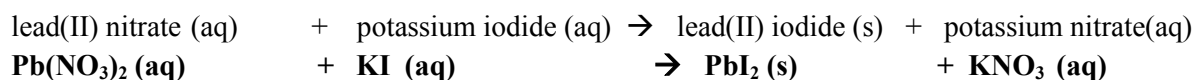
CHEMISTRY – NET IONIC EQUATION PRACTICE

Directions: For each chemical equation...

Write the formulas for the reactants and products in the unbalanced equation.

Write the balanced net ionic equation. You do not need to show all your work, only what is shown in boldface below.

EXAMPLE:



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- 1) sodium sulfate (aq) + barium nitrate (aq) \rightarrow sodium nitrate (aq) + barium sulfate (s)
 - 2) strontium oxide (aq) + aluminum perchlorate (aq) \rightarrow
strontium perchlorate (aq) + aluminum oxide (s)
 - 3) nitric acid (aq) + potassium hydroxide (aq) \rightarrow water (l) + potassium nitrate (aq)
 - 4) iron(III) bromide (aq) + silver nitrate (aq) \rightarrow iron(III) nitrate (aq) + silver bromide (s)
 - 5) potassium sulfide (aq) + calcium acetate (aq) \rightarrow potassium acetate (aq) + calcium sulfide (s)
 - 6) hydrochloric acid + ammonium hydroxide (aq) \rightarrow water (l) + ammonium chloride (aq)
 - 7) copper(II) chlorate (aq) + ammonium phosphate (aq) \rightarrow
copper(II) phosphate (s) + ammonium chlorate (aq)
 - 8) iron(II) acetate (aq) + sodium phosphate (aq) \rightarrow
iron(II) phosphate (s) + sodium acetate (aq)
 - 9) phosphoric acid (aq) + sodium hydroxide (aq) \rightarrow water (l) + sodium phosphate (aq)
 - 10) magnesium chloride (aq) + lithium carbonate (aq) \rightarrow
lithium chloride (aq) + magnesium carbonate (s)