

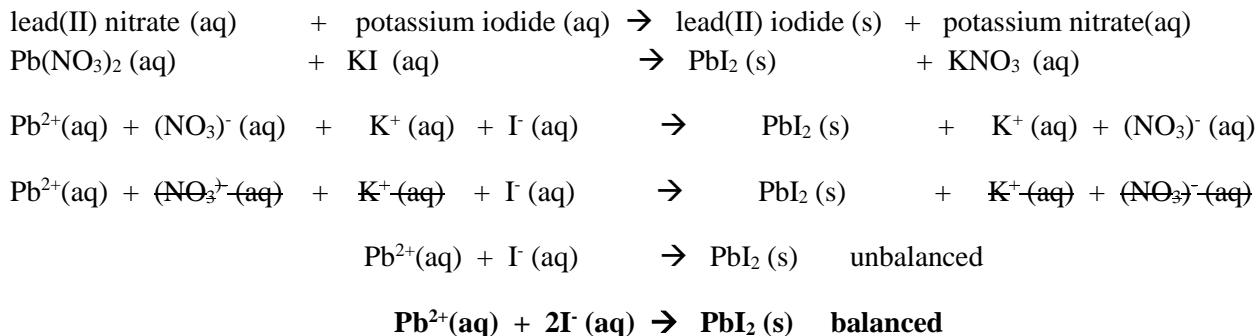
## CHEMISTRY – NET IONIC EQUATION PRACTICE (#4)

Directions: For each chemical equation...

Write the formulas for the reactants and products in the unbalanced equation.

Write the balanced net ionic equation. You do not need to show all your work, only what is shown in boldface below.

EXAMPLE:



- 1) sodium sulfate (aq) + barium nitrate (aq)  $\rightarrow$  sodium nitrate (aq) + barium sulfate (s)
- 2) strontium oxide (aq) + aluminum perchlorate (aq)  $\rightarrow$   
strontium perchlorate (aq) + aluminum oxide (s)
- 3) nitric acid (aq) + potassium hydroxide (aq)  $\rightarrow$  water (l) + potassium nitrate (aq)
- 4) iron(III) bromide (aq) + silver nitrate (aq)  $\rightarrow$  iron(III) nitrate (aq) + silver bromide (s)
- 5) potassium sulfide (aq) + calcium acetate (aq)  $\rightarrow$  potassium acetate (aq) + calcium sulfide (s)
- 6) hydrochloric acid (aq) + ammonium hydroxide (aq)  $\rightarrow$  water (l) + ammonium chloride (aq)
- 7) copper(II) chlorate (aq) + ammonium phosphate (aq)  $\rightarrow$   
copper(II) phosphate (s) + ammonium chlorate (aq)
- 8) iron(II) acetate (aq) + sodium phosphate (aq)  $\rightarrow$   
iron(II) phosphate (s) + sodium acetate (aq)
- 9) phosphoric acid (aq) + sodium hydroxide (aq)  $\rightarrow$  water (l) + sodium phosphate (aq)
- 10) magnesium chloride (aq) + lithium carbonate (aq)  $\rightarrow$   
lithium chloride (aq) + magnesium carbonate (s)