CHEMISTRY EMPIRICAL & MOLECULAR FORMULA PRACTICE

Show all work. Circle all answers. Show units and watch sig.figs.

EMPIRICAL FMLA.

1) What is the empirical formula for a compound which is 80.0% carbon and 20.0% hydrogen?

2) Calculate the empirical formula of a compound with 42.9% C and 57.1 % O.

MOLECULAR FMLA.

3) Find the molecular formula of a compound if the GFM is 30.0 g/mol and the empirical formula is the same as question #1.

4) What is the molecular formula of a compound with empirical formula of $C_3H_5O_2$ and a molar mass of 146.0 g/mol?

5) A compound has an empirical formula of $ClCH_2$ and a molecular mass of 98.96 g/mol. What is its molecular formula?

CHEMISTRY EMPIRICAL & MOLECULAR FORMULA PRACTICE

Show all work. Circle all answers. Show units and watch sig.figs.

EMPIRICAL FMLA.

1) What is the empirical formula for a compound which is 80.0% carbon and 20.0% hydrogen?

2) Calculate the empirical formula of a compound with 42.9% C and 57.1 % O.

MOLECULAR FMLA.

3) Find the molecular formula of a compound if the GFM is 30.0 g/mol and the empirical formula is the same as question #1.

4) What is the molecular formula of a compound with empirical formula of $C_3H_5O_2$ and a molar mass of 146.0 g/mol?

5) A compound has an empirical formula of $ClCH_2$ and a molecular mass of 98.96 g/mol. What is its molecular formula?