

## CHEM Electron Configuration Practice for IONS – Practice #2

### DIRECTIONS

- a) Name the *atom*.
- b) Using the periodic table, write the *VALENCE electron configuration for the neutral atom*.
- c) Write the *symbol and charge* of the ion that the element can form.
- d) Describe how many *electrons* will be *gained* or *lost* to form the ion.
- e) Write the *name* of the ion formed.
- f) Using the periodic table, write the *VALENCE electron configuration for the ion*.
- g) Give the symbol of the *Noble Gas* with which the ion is *isoelectronic*.

EXAMPLE: Al

- a) aluminum
- b)  $3s^2 3p^1$
- c)  $Al^{+3}$
- d) lose 3
- e) aluminum
- f) (third shell now empty,  $3s^0 3p^0$ )  $2s^2 2p^6$
- g) Ne

### QUESTIONS

- |        |
|--------|
| 1) Fr  |
| 2) Rb  |
| 3) Ba  |
| 4) O   |
| 5) Zn  |
| 6) S   |
| 7) Cd  |
| 8) Ni  |
| 9) Br  |
| 10) Sr |