

CHEMISTRY FORMULA PRACTICE #2

Write formulas for the following and *classify* as binary ionic (BI), binary molecular (BM), ternary ionic (TI), or other.

- 1) potassium nitride
- 2) potassium nitrate
- 3) potassium nitrite
- 4) aluminum hydroxide
- 5) diphosphorus trioxide
- 6) strontium cyanide
- 7) ammonium fluoride
- 8) lead(II) iodide
- 9) tin(IV) sulfide
- 10) sodium phosphate
- 11) iron(II) chloride
- 12) iron(III) perchlorate
- 13) copper(I) hypochlorite
- 14) zinc permanganate
- 15) calcium phosphide

Name the following and *classify* as binary ionic (BI), binary molecular (BM), and ternary ionic (TI), or other.

- 16) NaF
- 17) $\text{Ca}(\text{NO}_2)_2$
- 18) $\text{K}_2\text{Cr}_2\text{O}_7$
- 19) SCl_6
- 20) Na_3PO_3
- 21) As_2Br_5
- 22) NO_2
- 23) SnO
- 24) $(\text{NH}_4)_3\text{P}$
- 25) $\text{Al}_2(\text{SO}_3)_3$
- 26) CaCrO_4
- 27) $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$
- 28) $\text{Cr}(\text{HCO}_3)_3$
- 29) SO_3
- 30) Na_2Se