

Bauck's CHEM Ch. 13 Test Review

This is an optional assignment due the day of the test.

Materials: loose leaf paper, pen and/or pencil, calculator (You will be given a periodic table.)
Format: multiple choice; naming hydrates; math problems; short answer essays
Test date: _____
Test value: 200 points

VOCABULARY

- 1) **Colloid**—what is it; compare and contrast with **solution** and **suspension**
- 2) **Concentrated**—what is this; contrast to **dilute**
- 3) **Density** of water: liquid vs. solid (ice) – how do they differ?
- 4) **Dilute**—what is this; contrast to **concentrated**
- 5) **Electrolytes**—what are they; characteristics
- 6) **Hydrates**—what are they; how do their formulas look; be able to name hydrates
- 7) **“Like Dissolves Like”**—explain what this means and how it relates to forming solutions; give examples
- 8) **Molarity**—what does this measure; what is the molarity symbol and unit; be able to do molarity math problems (see next section below)
- 9) **Nonpolar**—what is it; give examples; contrast with **polar**
- 10) **Polar**—what is it; contrast with **nonpolar**
- 11) **Saturated** solution—what is it; characteristics; contrast with **unsaturated** and **supersaturated** solutions
- 12) **Solute**—what is it; give examples
- 13) **Solution**—what is it; give examples; compare and contrast with **suspension** and **colloid**
- 14) **Solvent**—what is it; give examples
- 15) **Specific heat**—what is it; how does it work; give examples of how high and low specific heat substances behave
- 16) **Supersaturated** solution—what is it; characteristics; contrast with **unsaturated** and **saturated** solutions
- 17) **Suspension**—what is it; compare and contrast with **solution** and **colloid**
- 18) **Unsaturated** solution—what is it; characteristics; contrast with **saturated** and **supersaturated** solutions
- 19) **Water**—characteristics; why is it a good solvent

MATH PROBLEMS

M = mol/L given on test

- You have to know how to use molar mass to go from g to mol.
 - For this review, give an example of each of the following. Show all work and units.
 - Be able to solve for molarity...
- 20) ...given moles and L
 - 21) ...given moles and mL (you have to know 1000 mL = 1L)
 - 22) ... given grams and L
 - 23) ... given grams and mL (you have to know 1000 mL = 1L)