

Chapter 12, part 1 review (simple mole calculations – pre-stoich)

- 1) How many L are in a mole of gas at STP?
- 2) Describe how to find the number of grams in one mole of substance.
- 3) How many representative particles (r.p.) are in one mole?
- 4) List the four types of r.p. and give an example of each.
- 5) How should all molar masses be rounded for this class?
- 6) What is STP?
- 7) What is D.A.?
- 8) What is the “teacher number”? Why is it important to D.A. problems?

For the following questions, set up a sample D.A.

- 9) Describe how you set up a problem converting liters to moles (for a gas at STP).
- 10) Describe how you set up a problem converting liters to grams (for a gas at STP).
- 11) Describe how you set up a problem converting moles to liters (for a gas at STP).
- 12) Describe how you set up a problem converting liters to r.p.
- 13) Describe how you set up a problem converting grams to moles.
- 14) Describe how you set up a problem converting r.p. to grams.
- 15) Describe how you set up a problem converting grams to liters (for a gas at STP).
- 16) Describe how you set up a problem converting moles to grams.
- 17) Describe how you set up a problem converting moles to r.p.
- 18) Describe how you set up a problem converting grams to r.p.
- 19) Describe how you set up a problem converting r.p. to moles.
- 20) Describe how you set up a problem converting r.p. to liters (for a gas at STP).