

Chem I H - ACID FORMULA WRITING AND NAMING – Practice #2

For #1-10:

- a) Name the following acids.
- b) Show what ions they contain (dissociate the formula – does not have to be balanced). See page 2 of the chemical reference sheet.

EXAMPLE: HCl

- a) hydrochloric acid
- b) $\text{HCl (aq)} \rightarrow \text{H}^{+1} \text{ (aq)} + \text{Cl}^{-1} \text{ (aq)}$

- | | |
|--------------------|-------------------------------------|
| 1) HBr | 6) HClO_3 |
| 2) HClO | 7) $\text{H}_2\text{C}_2\text{O}_4$ |
| 3) HI | 8) HNO_3 |
| 4) HClO_2 | 9) HSCN |
| 5) HF | 10) H_2CO_3 |

For #11-20:

- a) Write formulas for the following acids.
- b) Show what ions they contain (dissociate the formula – does not have to be balanced). See page 2 of the chemical reference sheet.

- 11) phosphoric acid
- 12) nitrous acid
- 13) acetic acid #1 AND acetic acid #2
- 14) phosphorous acid
- 15) carbonous acid
- 16) chromic acid
- 17) hydrosulfuric acid
- 18) perchloric acid
- 19) bromic acid
- 20) sulfuric acid