Pre-Stoich Mole In-Class Review

1) How many L are in a mole of gas at STP?
2) Describe how to find the number of grams in one mole of substance.
3) How many representative particles (r.p.) are in one mole?
4) List the four types of r.p. and give an example of each.
5) How should all molar masses be rounded for this class?
6) What is STP?
7) What is D.A.?
8) What is the "teacher number"? Why is it important to D.A. problems?

For the following questions, set up a sample D.A. ……..

9) Describe how you set up a problem converting liters to moles (for a gas at STP).
10) Describe how you set up a problem converting liters to grams (for a gas at STP).
11) Describe how you set up a problem converting moles to liters (for a gas at STP).
12) Describe how you set up a problem converting liters to r.p. (for a gas at STP).
13) Describe how you set up a problem converting grams to moles.
14) Describe how you set up a problem converting r.p. to grams.
15) Describe how you set up a problem converting grams to liters (for a gas at STP).
16) Describe how you set up a problem converting moles to grams.
17) Describe how you set up a problem converting moles to r.p.
18) Describe how you set up a problem converting grams to r.p.
19) Describe how you set up a problem converting r.p. to moles.
20) Describe how you set up a problem converting r.p. to liters (for a gas at STP).

Example for #9:

$$(\text{teacher \#})\ \cancel{\text{L}} \times \frac{1\ \text{mol}}{22.4\ \cancel{\text{L}}} = (\text{answer})\ \text{mol}$$