Bauck's CHEMISTRY Ch. 2 Test Review

This is an optional assignment due the day of the test.

Materials: loose leaf paper, pencil, calculator (clear memory if applicable)
Test date: _____
Test value: 200 points
Test format: (multiple choice), vocabulary application…
short answers, mole math, physical vs. chemical properties, physical vs. chemical changes, homogeneous vs. heterogeneous mixtures, mole math
-- If there is no multiple choice, most vocabulary will be on the midterm exam and will occur in other sections of the course.
-- Vocabulary with * will be needed for the rest of the test.

Topics:

1) **Alloy—** What is it? Is it chemical or physical? Give two examples.
2) ***Chemical change**—What is it? Give three examples. Contrast with physical change. (Be able to identify examples on the test).
3) **Chemical equation/reaction**— What is it? (Be able to identify on which side the **reactants** and **products** occur in an equation.)
4) ***Chemical property**—What is it? Give three examples. Contrast with physical property. (Be able to identify examples on the test).
5) **Chromatography—** What is it?
6) ***Compound—** What is it? Give three example formulas. (Be able to identify examples on the test).
7) ***Element—** What is this? Give three example formulas. (Be able to identify examples on the test).
8) **Filtration—** What is it? How does it work?
9) **Formula**— What is it? Be able to identify the number of atoms in a formula.
How many total atoms of C are in the formula $Zn(C_2H_3O_2)_2$?
10) ***Heterogeneous mixture**— What is it? Contrast with homogeneous mixtures. Give three examples for this review. (Be able to identify examples on the test).
11) ***Homogeneous mixture**— What is it? Contrast with heterogeneous mixtures. Give three examples for this review. (Be able to identify examples on the test).
12) ***Law of Conservation of Mass**— Summarize this law on your own words.
13) ***Physical change**—What is it? Give three examples. Contrast with chemical change. (Be able to identify examples on the test).
14) ***Physical property**—What is it? Contrast with chemical property. (Be able to identify examples on the test).
15) ***Product**— What is it? On which side of a chemical equation are the products found?
16) ***Reactant**— What is it? On which side of a chemical equation are the reactants found?
17) **Solute—** What is it? (Be able to identify examples on the test).
18) **Solvent**— What is it? (Be able to identify examples on the test).
19) **Solution**— What is it? (Be able to identify examples on the test).
20) ***Substance**—What is it? (Be able to identify examples on the test).

*** Note: There will be at least one question pertaining to material in past chapter(s) or unit(s). ***

MOLE MATH!