**Bauck's CHEMISTRY Ch. 1 part 2 Test Review**
*This is an optional assignment due the day of the test.*

Materials: loose leaf paper, pencil, calculator (clear memory if applicable)
Test date: _____
Test value: 200 points
Test format: all math problems and calculations…
metric conversions (practices 1-2), density (practice 3), sig.figs. (practice 4), mole D.A. (practice 5), general D.A. (practice 5)

| Equations you will be given: | $K = C + 273.15$ | $D = M/V$ |
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Topics:
1) **Absolute zero**— What is it? What is its significance?
2) **Accuracy**—What is it? Compare and contrast with **precision**.
3) **Base unit**— What is it? Give examples for the base units of length, mass, time, temperature, and amount of substance.
4) **Conversion factor**—What is it? How are they used in calculations? Give an example of a common conversion factor.
5) **Density**— What does this measure? (Be able to solve the density equation for D, M, or V.)
6) **Derived unit**— What is it? Compare and contrast with **base unit**. Give three examples from the notes.
7) **Mass**—What does this measure? Give three examples of mass units. (Be able to recognize and work with various mass units.)
8) **Metrics**—What do the following abbreviations mean?:  k, h, da, d, c, m (Be able to convert from one metric prefix to another.)
9) **Scientific notation**—How does this work? Give an example.
10) **Significant figures** ("Sig.Figs.")—Why are these used in science calculations? (Be able to determine the number of sig.figs and calculate problems to the correct amount of sig.figs.)
11) **Temperature**—What does it measure? Give three examples of temperature scales from the notes. (Be able to recognize and work with various temperature units.)
12) **Volume**—What does this measure? Give three examples of volume units. (Be able to recognize and work with various volume units.)
13) Math Problems… For this review, give an example of a solved math problem for each of the following:
    a. Density
    b. Dimensional Analysis with time, distance, or metrics
    c. (Mole → gram)  or  (gram → mol) DA
    d. (Mole → liter)  or  (liter → mol) DA
    e. (Mole → particle)  or (particle → mol) DA
    f. Metric prefix conversions
    g. Scientific notation
    h. Significant figures
    i. Temperature conversions using $K = C + 273$

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*** Note:  There will be at least one question pertaining to material in past chapter(s) or unit(s).  ***